

NAME \_\_\_\_\_

SECTION \_\_\_\_\_ DATE \_\_\_\_\_

INSTRUCTOR \_\_\_\_\_

## EXERCISE 3

### Names and Formulas I

Give the names of the following compounds:

1. NaCl

\_\_\_\_\_

2. AgNO<sub>3</sub>

\_\_\_\_\_

3. BaCrO<sub>4</sub>

\_\_\_\_\_

4. Ca(OH)<sub>2</sub>

\_\_\_\_\_

5. ZnCO<sub>3</sub>

\_\_\_\_\_

6. Na<sub>2</sub>SO<sub>4</sub>

\_\_\_\_\_

7. Al<sub>2</sub>O<sub>3</sub>

\_\_\_\_\_

8. CdBr<sub>2</sub>

\_\_\_\_\_

9. KNO<sub>2</sub>

\_\_\_\_\_

10. Fe(NO<sub>3</sub>)<sub>3</sub>

(a)

\_\_\_\_\_

(b)

\_\_\_\_\_

11. (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub>

\_\_\_\_\_

12. KClO<sub>3</sub>

\_\_\_\_\_

13. MgS

\_\_\_\_\_

14. Cu<sub>2</sub>C<sub>2</sub>O<sub>4</sub>

\_\_\_\_\_



NAME \_\_\_\_\_

SECTION \_\_\_\_\_ DATE \_\_\_\_\_

INSTRUCTOR \_\_\_\_\_

## EXERCISE 4

### Names and Formulas II

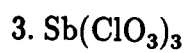
Give the names of the following compounds:



\_\_\_\_\_



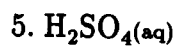
\_\_\_\_\_



\_\_\_\_\_



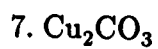
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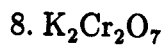


(a)

\_\_\_\_\_

(b)

\_\_\_\_\_



\_\_\_\_\_

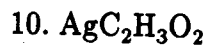


(a)

\_\_\_\_\_

(b)

\_\_\_\_\_



\_\_\_\_\_



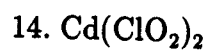
\_\_\_\_\_



\_\_\_\_\_



\_\_\_\_\_



\_\_\_\_\_

Give the formulas of the following compounds:

1. Sodium oxalate 1. \_\_\_\_\_
2. Manganese(II) iodate 2. \_\_\_\_\_
3. Zinc nitrite 3. \_\_\_\_\_
4. Potassium permanganate 4. \_\_\_\_\_
5. Titanium(IV) bromide 5. \_\_\_\_\_
6. Sodium arsenate 6. \_\_\_\_\_
7. Manganese(IV) sulfide 7. \_\_\_\_\_
8. Bismuth(III) arsenate 8. \_\_\_\_\_
9. Sodium peroxide 9. \_\_\_\_\_
10. Magnesium bicarbonate 10. \_\_\_\_\_
11. Lead(II) acetate 11. \_\_\_\_\_
12. Phosphoric acid 12. \_\_\_\_\_
13. Nitric acid 13. \_\_\_\_\_
14. Acetic acid 14. \_\_\_\_\_
15. Arsenic(III) iodide 15. \_\_\_\_\_
16. Ammonium thiocyanate 16. \_\_\_\_\_
17. Cobalt(II) chlorite 17. \_\_\_\_\_
18. Stannous fluoride 18. \_\_\_\_\_

NAME \_\_\_\_\_

SECTION \_\_\_\_\_ DATE \_\_\_\_\_

INSTRUCTOR \_\_\_\_\_

## EXERCISE 5

### Names and Formulas III

Give the names of the following compounds:

1.  $\text{CO}_2$  \_\_\_\_\_
2.  $\text{H}_2\text{O}_2$  \_\_\_\_\_
3.  $\text{Ni}(\text{MnO}_4)_2$  \_\_\_\_\_
4.  $\text{Co}_3(\text{AsO}_4)_2$  \_\_\_\_\_
5. KCN \_\_\_\_\_
6.  $\text{Sb}_2\text{O}_5$  \_\_\_\_\_
7.  $\text{BaH}_2$  \_\_\_\_\_
8.  $\text{NaHSO}_3$  \_\_\_\_\_
9.  $\text{As}(\text{NO}_2)_5$  \_\_\_\_\_
10. KSCN \_\_\_\_\_
11.  $\text{Ag}_2\text{CO}_3$  \_\_\_\_\_
12.  $\text{CrF}_3$  \_\_\_\_\_
13.  $\text{SnS}_2$  (a) \_\_\_\_\_  
(b) \_\_\_\_\_
14.  $\text{H}_2\text{SO}_3(\text{aq})$  \_\_\_\_\_
15.  $\text{HgC}_2\text{O}_4$  \_\_\_\_\_
16.  $\text{Pb}(\text{HCO}_3)_2$  \_\_\_\_\_
17.  $\text{Cu}(\text{OH})_2$  \_\_\_\_\_

## EXERCISE 5 (continued)

NAME \_\_\_\_\_

KEY \_\_\_\_\_

Give the formulas of the following substances:

- |                                |           |
|--------------------------------|-----------|
| 1. Ammonium hydrogen carbonate | 1. _____  |
| 2. Hydrogen sulfide            | 2. _____  |
| 3. Barium hydroxide            | 3. _____  |
| 4. Carbon tetrachloride        | 4. _____  |
| 5. Nickel(II) perchlorate      | 5. _____  |
| 6. Lead(II) nitrate            | 6. _____  |
| 7. Sulfur dioxide              | 7. _____  |
| 8. Carbonic acid               | 8. _____  |
| 9. Copper(I) carbonate         | 9. _____  |
| 10. Calcium cyanide            | 10. _____ |
| 11. Arsenic(III) oxide         | 11. _____ |
| 12. Silver dichromate          | 12. _____ |
| 13. Nitrous acid               | 13. _____ |
| 14. Copper(II) bromide         | 14. _____ |
| 15. Ammonia                    | 15. _____ |
| 16. Chlorine                   | 16. _____ |
| 17. Chromium(III) sulfite      | 17. _____ |
| 18. Chloric acid               | 18. _____ |
| 19. Barium arsenate            | 19. _____ |
| 20. Manganese(IV) chloride     | 20. _____ |
| 21. Carbon disulfide           | 21. _____ |
| 22. Aluminum fluoride          | 22. _____ |

43. The oxidation number of manganese in  $\text{KMnO}_4$  is \_\_\_\_\_.

44. The oxidation number of antimony in  $\text{Sb}_2\text{S}_5$  is \_\_\_\_\_.

### Name the Compound Represented by the Formula

45.  $\text{CdS}$  \_\_\_\_\_

46.  $\text{CaC}_2$  \_\_\_\_\_

47.  $\text{BCl}_3$  \_\_\_\_\_

48.  $\text{NH}_3$  \_\_\_\_\_

49.  $\text{MnO}_2$  \_\_\_\_\_

50.  $\text{XeF}_4$  \_\_\_\_\_

51.  $\text{NaNO}_3$  \_\_\_\_\_

52.  $\text{Hg}_2\text{Cl}_2$  \_\_\_\_\_

53.  $\text{ZnCO}_3$  \_\_\_\_\_

54.  $\text{Na}_2\text{Cr}_2\text{O}_7$  \_\_\_\_\_

55.  $\text{PbSO}_4$  \_\_\_\_\_

56.  $\text{NH}_4\text{NO}_3$  \_\_\_\_\_

For more practice try text Exercises 34–38.

### Write Formulas

57. Aluminum phosphate \_\_\_\_\_

58. Cobalt(II) nitrate \_\_\_\_\_

59. Barium peroxide \_\_\_\_\_

60. Barium chloride \_\_\_\_\_

61. Chromium(III) sulfide \_\_\_\_\_

62. Lithium hydroxide \_\_\_\_\_

63. Mercury(II) oxide \_\_\_\_\_

64. Iodine heptafluoride \_\_\_\_\_

65. Ammonium sulfate \_\_\_\_\_

66. Barium hydroxide \_\_\_\_\_

67. Potassium hydride \_\_\_\_\_

68. Mercury(I) nitrate \_\_\_\_\_

For more practice try text Exercises 40 and 41.